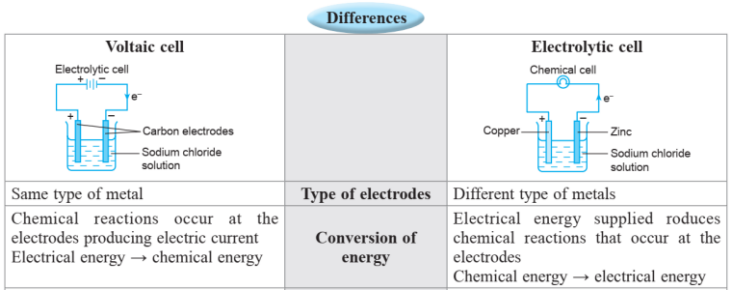
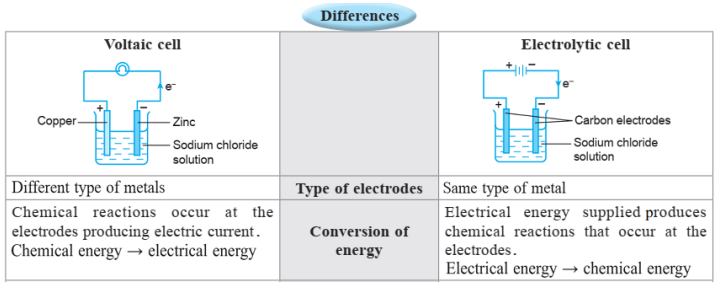
Title: Focus SPM Chemistry 2021
Book Code: CC038641
Author: Chien Hui Siong, Low Swee Neo, Lim Eng Wah, Salida Sani

Errata

Page number	Section / Part	Error	Correction
Page 155	**Checkpoint 6.8**	**Q2** (b) Phosphoric acid + Calcium hydroxide	**Q2** (b) Phosphoric acid + Potassium hydroxide
Page 273	**Comparison between Electrolytic Cell and Voltaic Cell**		
Page 293	**Checkpoint**	**S4** **S5** **S6**	**Q4** **Q5** **Q6**
Page 297	**2 (Figure 2.7)**	**Akhiran**	**Suffix**
	3	Name it appropriately as meth-, eth-, prop-, but-, pent-, etc (Table 2.5)	Name it appropriately as meth-, eth-, prop-, but-, pent-, etc (Table 2.7)
Page 306	**Table 2.15**	5th column: **Solubility (g per 100 g of water)** Bercampur dalam semua kadaran Bercampur dalam semua kadaran Bercampur dalam semua kadaran	5th column: **Solubility (g per 100 g of water)** Miscible in all proportions Miscible in all proportions Miscible in all proportions
Page 311	**Figure 2.15**	Misplaced of 'Alkene' with 'Alkane'	Exchange label of 'Alkene' with 'Alkane' or vice versa
Page 338	**Table 2.21**	2nd column:	2nd column:

<table>
<tr><td></td><td></td><td>Molecular formula, C_nH_{2n-2}
4
5</td><td>Molecular formula, C_nH_{2n-2}
C_4H_6
C_5H_8</td></tr>
<tr><td>Page 353</td><td>Results</td><td>
<table>
<tr><th>Mixture</th><th>Initial temperature (°C)</th><th>Highest/ lowest temperature (°C)</th><th>Heat is released or absorbed</th><th>Type of reaction</th></tr>
<tr><td>Sodium hydroxide + water</td><td>33</td><td>39</td><td>Heat is released</td><td>Exothermic</td></tr>
<tr><td>Anhydrous calcium chloride + water</td><td>33</td><td>41</td><td>Heat is released</td><td>Exothermic</td></tr>
<tr><td>Ammonium chloride + water</td><td>33</td><td>29</td><td>Heat is absorbed</td><td>Exothermic</td></tr>
<tr><td>Sodium thiosulphate + water</td><td>33</td><td>31</td><td>Heat is absorbed</td><td>Exothermic</td></tr>
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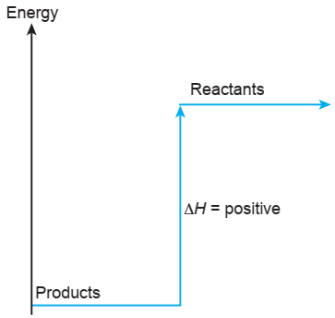
</table>
</td></tr>
<tr><td>Page 354</td><td>Figure 3.9</td><td> 

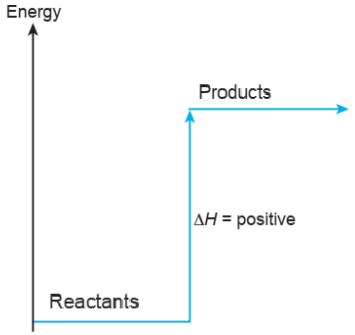
Figure 3.9 Energy level diagram for endothermic reaction</td><td> 

Figure 3.9 Energy level diagram for endothermic reaction</td></tr>
<tr><td>Page 381</td><td>Solution</td><td>$$\text{Number of moles of } HNO_3 = \frac{MV}{1000} = \frac{M \times 26}{1000} = 0.026 \text{ mol}$$</td><td>Number of moles of $HNO_3 = \frac{MV}{1000}$</td></tr>
</table>

			$= \frac{M \times 26}{1000}$ $= 0.026M$ mol